

has an anti-Andean direction, NE – SW, typical morphology of the southern Andes of Peru.



Figure 2: View of geothermal manifestations in Calientes Valley.

Calientes is characterized for the presence of several volcanic systems and tectonic structures with mainly direction NE-SW (See Fig N° 3). In the west part of the valley we can see the volcanic complex Yucamane – Calientes – Yucamane constituted principally by lava flows, domes of lava and some remnants of pyroclastic flows. Of all these volcanoes Yucamane present more recent activity, whereas Calientes and Yucamane Chico volcanoes, located more in the northern part presented activity possibly in before the Pleistocene (datation in process).

As a products expressed by the volcano Yucamane we distinguished principally flows of andesitic lava, blocks and ashes pyroclastic flows associate to domes of lava. The stratigraphy of the Yucamane volcano indicates a several volcanic events.

Holocenic deposits correspond to alluvial, colluvial and swamp, they are shaped by accumulations of gravels, sands and silt reworked, and they principally are refilling the current riverbeds of streams and rivers. In the middle and top part of the Calientes Valley we have a big sinter deposit, which is formed by the precipitation of carbonates and silica. The swamps are located mainly in the streams, they have saturated sediments.

In the southern part of Yucamane volcano, there are small outcrops of quartz sandstone from Hualhuani Formation, Yura Group. Likewise on the south of Geothermla Field of Calientes there are outcrops of limestones from Gramadal Formation, from Yura group also. With this information we are interpreted that sedimentary rocks are the bedrock on which there have settled the diverse volcanic episodes that happened during the Neógene.

In Calientes Geothermal Field there is a regional geological structure that is characterized by a structural fault which has NW-SE tendency and it is parallel to the Andean and volcanic chain direction. The volcanoes around the Calientes are aligned in direction NE-SW, and this indicate that there is a fault which this direction. The structural system in Calientes is summarized in this way: structures with NE-SW tendency, they are well identified in the field and goes along Calientes river which has a lot of hot spring with this direction. Therefore this system of lineaments seems to be the fault of the system (Fc-1). In addition to the described system, faults have been identified with N-S direction which determines the course of Calientes River (Fc-2) in the south part of the valley, along these structures there are not geothermic manifestations. Likewise the analysis of images satellites indicates the presence of fault with NNO-SSE direction (Fc-3) (See Fig. N° 4) (Vargas, 2007).

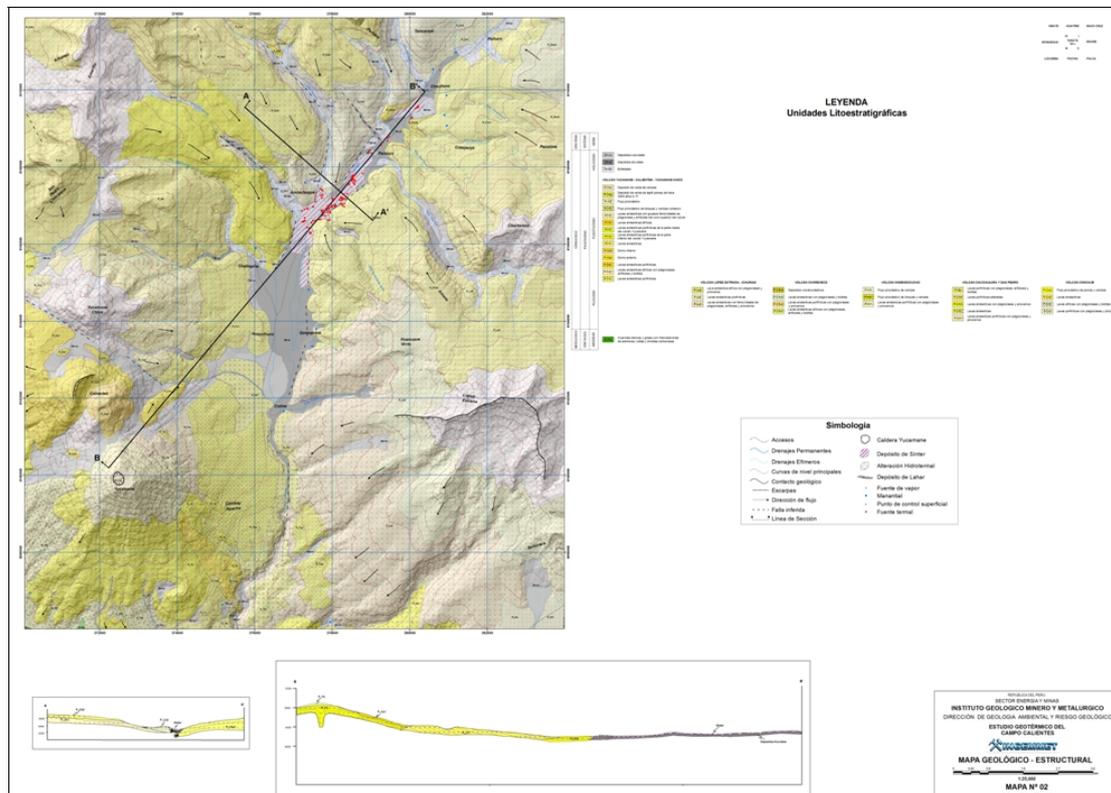


Figure 3: Calientes Geological Map.

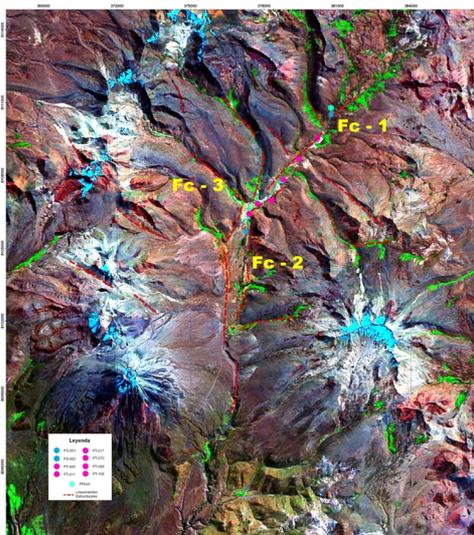


Figure 4: Mainly structural aligned in the Geothermal Field of Calientes.

3. CHEMICAL COMPOSITION OF THE GEOTHERMAL FLUID

3.1. Sampling and Analysis

Geochemical studies of geothermal fluids involve three main steps: 1) sample collection, 2) chemical analysis and 3) data interpretation.

3.1.1. Sampling

Six geothermal water samples with a temperature range of 50 °C to 90 °C were collected from geothermal wells in the Calientes field to analyses, also we have collected three cold waters samples from high and low part of the Calientes River, as well as of a cold source.

We take 3 samples: One for the analyses of cations, it was preserved with addition chloride acid; a second one is taken without addition of acid, to determine anions (non – preserved sample). A third sample, to determine isotopic composition. These samples have enabled us to establish a basis for explaining the chemistry of thermal water in the Calientes field.

The isotope analysis for deuterium (δD), oxygen-18 (^{18}O) and chemical composition of water samples was determined by the West Japan Engineering Consultants, INC in Japan. The different methods adopted are as listed in table 1.

3.1.2. Analysis on Site

Temperature

The measured of the temperature can be made with a mercury thermometer or a thermocouple. The precision is generally 0.1°C.

Conductivity

We measure the conductivity at the sampling temperature, with a conductimeter, but conductivity is function of the temperature (it decrease when temperature also decrease). So it's necessary to report all values of conductivity at a same temperature (usually 25°C) to allow comparisons between waters. The unit is Siemens (S).

pH

The value of pH also depends of the temperature. We use a glass-electrode pH-meter with a precision of 0.01.

3.1.3. Analytical Results

The results of the chemical analysis, together with the isotopic composition of water samples from springs in the Calientes field are listed in tables 1 and 2, respectively.

During the field work, carried out between October 10 and 11, 2007 (Table 1), we have note that the waters associates to Calientes Geothermal Field, present the maximum measured temperature was 87 °C and the minimum 50 °C and pH in the range 6 to 8, with exception of the water of the Calientes River in the high part that there presented an pH acid of 3.1 (Table 1).

The 6 thermal springs (TC-4, TC-10, TC-15, TC-31, TC-41, TC-47) directly associated with the Calientes Geothermal Field, have showed concentrations of sodium and potassium between ranges from 774mg/L to 1190mg/L and 48.6mg/L to 112 mg/L, respectively, also we have observed high concentrations of the chloride ion, whit ranges from 271mg/L to 1900 mg/L and boron from 32.1 mg/L to 54.8 mg/L. Some components are not present in the list in the following tables because their concentrations were very low, such as the case of the NO₃⁻, I⁻ and NO₂⁻.

The majority of these thermal springs present bubbles which indicates that there is presence of CO₂. Bicarbonate concentration ranges from 83mg/L to 195mg/L, the sulfates concentration from 21mg/L to 128mg/L respectively. The concentration of calcium and magnesium are much lower than of the other major anions. The chloride is the principal dominant anion in these waters, followed by the sodium. Also we have been observed that the waters are mainly of pH neutral and chemically are Na - Cl waters, typical of the deep geothermal fluid found in most high-temperature systems.

4. DISCUSSION

The Schoeller diagram (Marini, 2000) (Fig. 5) show that the springs TC-4, TC-10, TC-15, TC-41 and TC-47, presented inside of their chemical composition high concentrations of sodium (Na) and chloride (Cl), unlike the source TC-31 which values are low in Cl, SO₄ and HCO₃, possibly for suffering dilution with meteoric waters. Also we notice that the sample collected in the high part of the river, present as dominant anion the SO₄ typical of acid waters, coinciding with the pH of 3.1. Likewise, we notice that the sample collected in the low part of the river presents high concentration in chlorides, which indicates that the water of the Calientes River is mixing by waters with high concentrations of chlorides, from the thermal springs, which are located nearby to the shore.

The chemical results obtained in the thermal waters, also were plotted in the Langelier diagram (Fig. 6) (Fyticas et al., 1989), where we have observed that all the thermal waters gather in crowds inside the quadrant of the chemical proportions Cl + SO₄ - Na + K, specifically in the Na-Cl sector. This indicates us that the waters are of the type sodium -chloride ó alkaline, characteristic of geothermal waters. Likewise we notice that the spring TC-31 and the Calientes River water is separate, this is due to the fact that the waters are suffering a mixture with meteoric waters and in case Calientes River with the thermal waters.

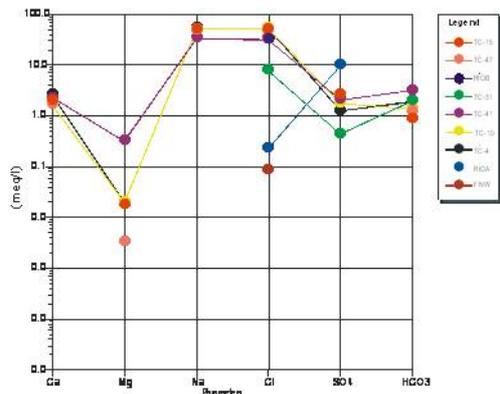


Figure 5: Schoeller diagram.

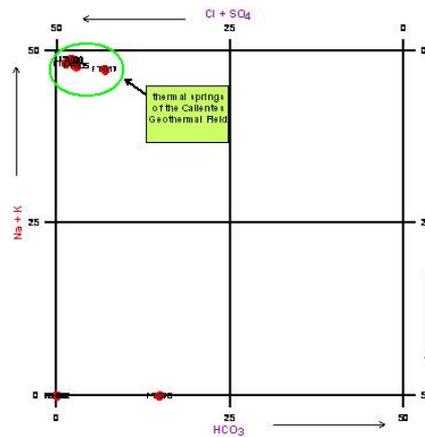


Figure 6: Langelier diagram.

According to the ternary diagram Cl-SO₄-HCO₃ (Giggenbach, 1988) (Fig. 7), the thermal springs are inside the deep chlorides waters group, that is typical of deep geothermal fluids in the most high-temperature system. Areas which contain hot, large-flow springs with the greatest Cl concentration are fed more directly from the deep reservoir, and identify permeable zones within the field. This wants to say that the waters in the Calientes Geothermal Field come from deep level reservoirs (Nicholson, 1993).

Following the Piper's diagram (Fig. 8), the waters in study locate also inside the quadrant alkaline chloride. These indicate that the chemical principal type of water geothermal is Na-Cl, being the dominant ions in the chemical composition of the waters. Piper's diagram does not do any more that to support the hypothesis showed by Langelier diagram.

The binary diagram B - Cl (Fig. 9) show that the geothermal waters in Calientes present high concentration of boron, that could be from the meteoric alteration of volcanic rocks and to the activity hydrothermal (Risacher, 1984) or can be to related to the meteoric leaching and/or the hydrothermalization of rocks rich in boron (Murray, 1996).

The geothermal waters in Calientes field shows high Cl/B mole ratio, typically of the neutral Na-Cl type hot water of low salinity from hot springs and the geothermal wells, as well as on the geologic structures and the volcano distributions around these areas (Shigeno, 1993). Namely the essential geothermal reservoir is developed in the marine sedimentary rocks, which probably have relatively

high porosity and permeability with abundant fractures (Shigeno and Abe 1983). In case of the Calientes area we have observed that the geothermal waters are located inside of the volcanic systems alienated of SW-NE, where the geology shows several faults and fractures that allowed the outcrop of hot waters, whose reservoir to deep is probably heated mainly by the thermal conduction from the magma chambers of the volcanoes located in this area. Likewise we might to indicate that the geothermal waters are high in B and B/Cl implying reaction with marine sedimentary rocks at deep level.

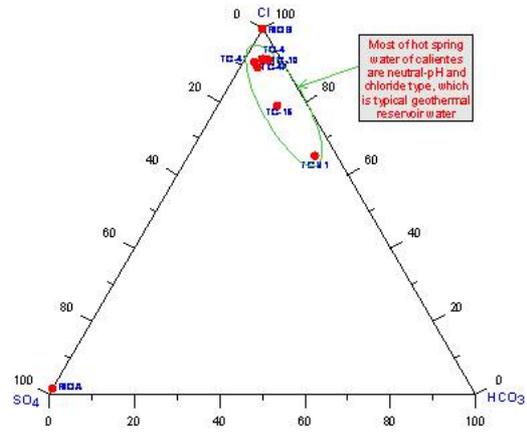
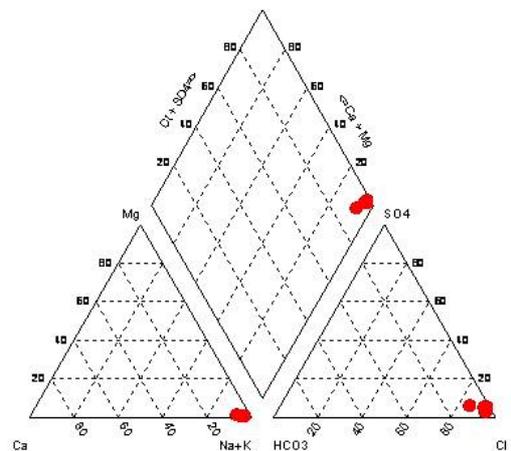


Figure 7: Ternary diagram (Giggenbach, 1998).



Figures 8: Piper diagram.

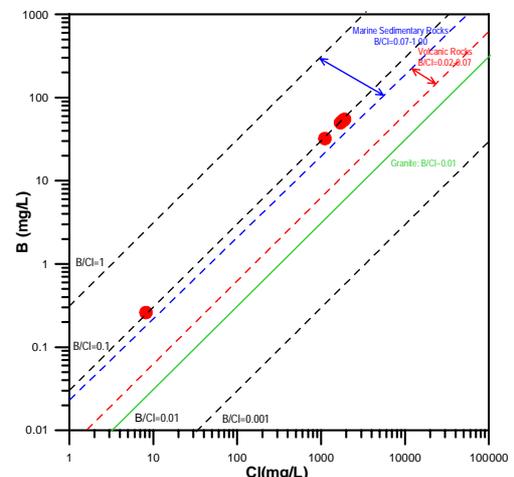


Figure 9: Boro vs Cloruro diagram (Shigeno, 1993).

The evaluation of the analytical data plotted in the ternary diagram Na-K-Mg (Giggenbach et al., 1983) (Fig.10) shows that the majority samples fall in the field of the equilibrium line characteristics of mature or geothermal waters, suggesting that their cation ratios are controlled by mineral-solution equilibrium. In case the spring TC-15 that fall in the field of the partial equilibrium is probably that suffer a dilution process or mixing with superficial waters.

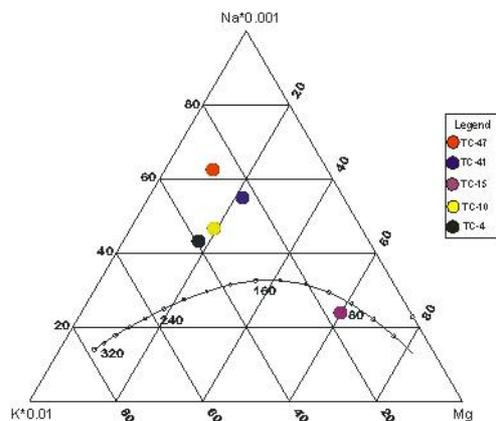


Figure 10: Ternary diagram (Giggenbach, 1983).

5. ISOTOPIC RESULTS

The water sample were analyzed to determined isotopic ratios of $\delta^{34}\text{S}$ (SO_4), $\delta^{18}\text{O}$ (SO_4), $\delta^{18}\text{O}$ and δD , the isotopic results (Table 2) of the thermal waters of Calientes shows that the cold waters present more negative values of $\delta^{18}\text{O}$ than geothermal waters, which is separate from the meteoric water line, but is not very higher, that indicate geothermal waters associates with magmatic fluids are mixing with meteoric waters (see Fig. 11) (Craig et to., 1956; Craig, 1963).

The Figure 11 shows the relationship between $\delta^{18}\text{O}$, ranging from -16.1 to -13.3‰, and δD , ranging -122 a -111‰. Two samples were plot close to the Global Meteoric Water Line (Craig, 1961), indicating that these waters are mainly fed by meteoric water. Therefore we might to indicate that the spring water of Calientes comprises 80% meteoric water and 20% magmatic water (see Fig. 11).

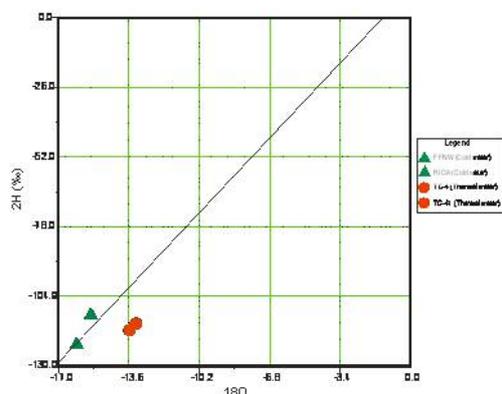


Figure 11: δD vs $\delta^{18}\text{O}$ in the Calientes Field.

The thermal waters are enriched both in $\delta^{18}\text{O}$ and δD isotopes. Interaction of the thermal waters with the surrounding rocks is evident in the temperature relationship for the isotopes (Hitchon and Friedman, 1969). Also we

notice clearly in the Figures 12 and 13 that the cold waters and geothermal waters shows the same values in δD isotopes, while the $\delta^{18}\text{O}$ values of the geothermal waters is more positive than cold waters. Likewise we have observed that the $\delta^{18}\text{O}$ isotopes in the geothermal waters increase with the temperature, which implies that the interaction between water and rocks is affected by temperature.

According White (1974) the composition of juvenile or magmatic waters have a range of isotope content, with $\delta^{18}\text{O} +6$ to $+9$ ‰ and $\delta\text{D} -40$ to -80 ‰, in case of the waters in study the δD value is inside the mentioned values by White, but for the $\delta^{18}\text{O}$ is outside of the values, probably, this can be then further changed by several sub-surface processes as the geothermal water rises to the surface.

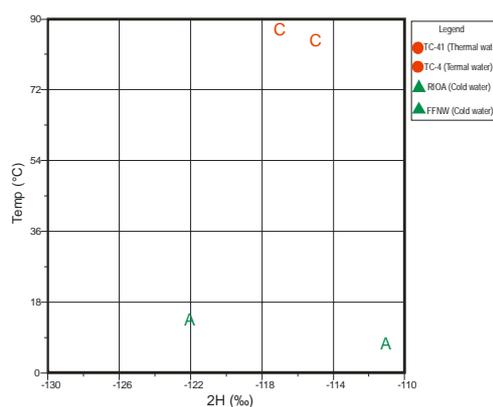


Figure 12: Temperature vs δD .

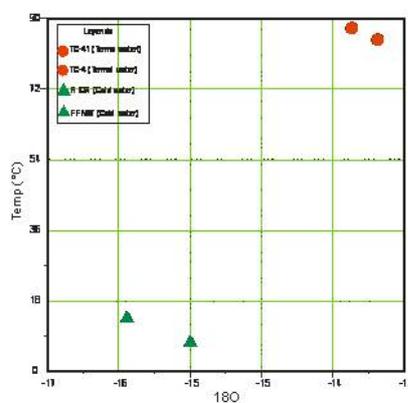


Figure 13: Temperature vs $\delta^{18}\text{O}$.

6. GETHERMOMETRY RESULTS

The purpose of geothermometer is estimate the deep temperature, given concentrations of dissolved substances at the surface. This method assumes that concentrations at depth are preserved as the waters flow to the surface, and measures the degree to which the substances were in equilibrium at depth. Geothermometer calculations are empirical, but in most cases seem to give a reasonable estimate.

The chemical geothermometer results calculated for the thermal water in the Calientes Geothermal Field are presented in Table 3, which in liquid phase allowed us to estimate the temperature to depth of the geothermal resource in the Calientes Geothermal Field. According to the results obtained of the geothermometers (Table 3 and

Fig. 14), we have estimated that the temperature of the geothermal resource is bigger than 200 °C (between 240 °C and 280 °C) to 2000 m of depth approximately.

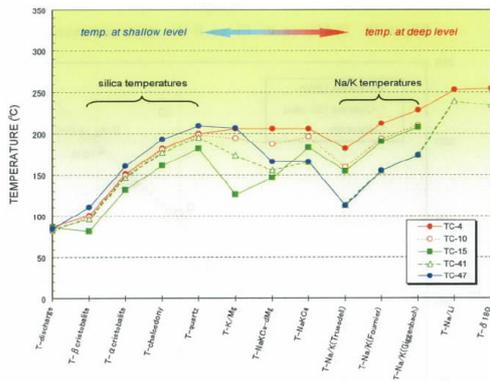


Figure 14: Geothermometer temperatures for the thermal waters in the Calientes field.

7. CONCLUSION

- The geochemical characterization of the thermal springs in Calientes are classified as Alkaline-chloride (Na-Cl).

- Stable isotopes $\delta^{18}O$ and δD relationship proves that the thermal waters are coming up from a mixing of meteoric and magmatic waters.

- For estimation of aquifer temperature of the geothermal resource is bigger than 200°C (between 240 and 280°C) to 2000 m of depth approximately.

- According to the geothermal manifestation and estimated subsurface temperature by geochemical thermometers, the geothermal resources in this field seem to be promising. However, data and information on characteristics and structures of geothermal resources are not enough to make geothermal reservoir model and to evaluate generation power capacity of this field. Geoscientific studies for clarification of characteristics and structures of the geothermal resources had been necessary as resource feasibility study. After the geoscientific surface study, exploratory well drilling and well test should be conducted to confirm geothermal resource existence and to evaluate its capacity.

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Table 1: Chemical composition of the thermal waters in Calientes (WEST JEC).

Source	Temp. °C	pH	Cond. ($\mu\text{S/cm}$)	Na mg/L	K mg/L	Ca mg/L	Mg mg/L	Cl mg/L	SO ₄ mg/L	HCO ₃ mg/L	B mg/L	As mg/L	SiO ₂ mg/L
FFNW	26.9	7.24	110					3.05				0.084	
RIOA	7.4	3.40	907					8.17	479		0.26	<0.005	
TC-4	87.3	7.60	5,920	1190	112	52.2	0.235	1900	60.2	117	54.8	8.66	263
TC-10	81.7	7.44	5,900	1180	88.3	32.5	0.242	1860	84.1	83	53.4	6.33	258
TC-15	86.6	6.90	3,900	774	55.4	45.6	3.87	1110	99.1	195	32.1	5.01	204
TC-31	54.8	7.53	1,160					271	21	123			141
TC-41	84.3	7.12	5,690	1180	51.6	39.1	0.224	1810	128	53	52.7	7.88	248
TC-47	84.3	7.90	5,380	1120	48.6	33.5	0.043	1700	122	78	49.8	8.33	297
RIOB	23.8	8.34	4,140					1170				3.93	

Table 2: Isotopic results.

Source	Fecha	δ D(H ₂ O)	δ ¹⁸ O(H ₂ O)	δ ³⁴ S(SO ₄)	δ ¹⁸ O(SO ₄)
FFNW	10/10/2007	-122	-16.1		
RIOA	10/10/2007	-111	-15.4		
TC-4	10/10/2007	-117	-13.6	6.5	-7.4
TC-41	11/10/2007	-115	-13.3	10	-6.3

Table 3: Chemical geothermometer results.

Code	T- Qua- rzo (°C)	T- Calce- donia (°C)	T- α Cristob- alita (°C)	T- β Cristo- balita (°C)	T- Na/K (T) (°C)	T- NaKC a (°C)	T- NaKCa- Mg (°C)	T- K/Mg (°C)	T-Na/Li (°C)	T - δ ¹⁸ O(H ₂ O-SO ₄) (°C)
TC-4	199	181	151	101	181	228	205	205	252	254
TC-10	198	180	149	99	159	233	187	194		
TC-15	181	160	132	82	154	182	146	126		
TC-31	157	133	107	58						
TC-41	195	176	146	96	113	192	156	173	239	235
TC-47	209	192	160	110	113	193	165	206		